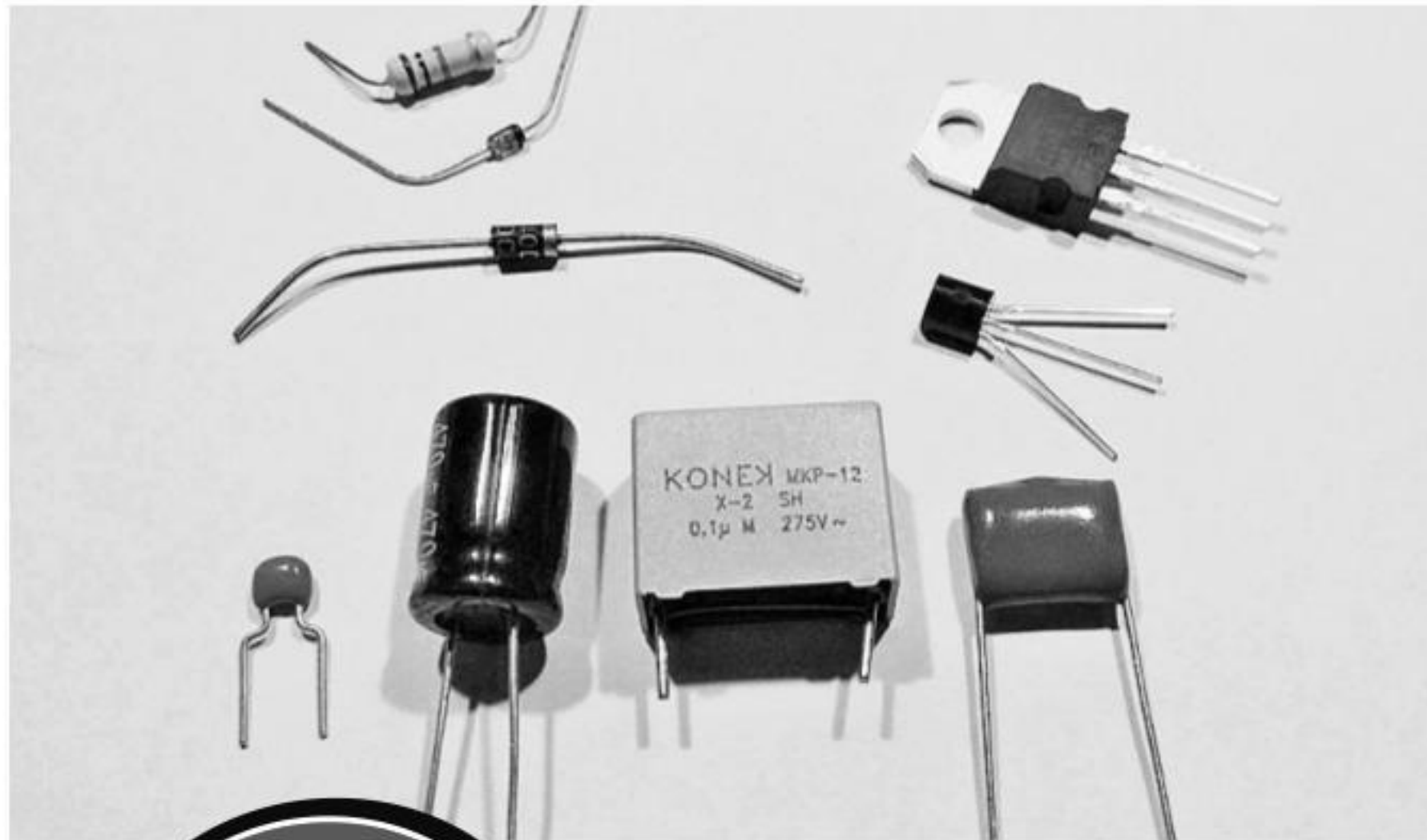


# Symbol for zener diode is..

## BASIC ELECTRONICS

- Copper is an example of
  - 1) Conductor
  - 2) Semiconductor
  - 3) Insulator
  - 4) none of these
- At  $0^{\circ}K$  semiconductor behaves like
  - 1) conductor
  - 2) semiconductor
  - 3) insulator
  - 4) none of these
- The forbidden energy gap for silicon is
  - 1) 0.3 eV
  - 2) 3.5 eV
  - 3) 0.7 eV
  - 4) 1.12 eV
- When a semiconductor is doped with a P-type impurity, each impurity atom will
  - 1) acquire negative charge
  - 2) acquire positive charge
  - 3) remain electrically neutral
  - 4) give away one electron
- In an intrinsic semiconductor
  - 1) there are no holes in the material
  - 2) the number of holes is too small
  - 3) no of electrons in CB equal to no. of holes in VB
  - 4) there are no electrons in the material
- N-type semiconductor is
  - 1) -ve ionised
  - 2) +ve ionised
  - 3) electrically neutral
  - 4) none of these
- In N-type semiconductor electrons are called
  - 1) majority charge carriers
  - 2) minority charge carriers
  - 3) both 1 and 2
  - 4) none
- In P-type semiconductor electrons are called
  - 1) majority charge carriers
  - 2) minority charge carriers
  - 3) none of these
  - 4) all of the above
- P-type semiconductor is an example of
  - 1) intrinsic
  - 2) pure
  - 3) impure
  - 4) none of these
- The process by which impurities are added to a pure semiconductor is called
  - 1) diffusion
  - 2) drift
  - 3) doping
  - 4) mixing
- PN-junction is formed by
  - 1) joining P-type and N-type together
  - 2) soldering P-type and N-type
  - 3) doping P-type from one side and N-type from the other side for pure semiconductor
  - 4) none of these
- The junction voltage for Ge diode is
  - 1) 0.6 V
  - 2) 0.2 V
  - 3) 0.1 V
  - 4) 0.01 V
- For an ideal diode forward based resistance is
  - 1) zero
  - 2) minimum
  - 3) maximum
  - 4) infinite
- For an ideal diode, the reverse biased resistance is
  - 1) zero
  - 2) minimum
  - 3) maximum
  - 4) infinite
- The diffusion current in diode is caused by
  - 1) voltage
  - 2) heat energy
  - 3) crystal formation
  - 4) electrical energy
- When diode is reverse biased, the barrier potential will
  - 1) increase
  - 2) decrease
  - 3) remain same
  - 4) none of these
- When diode is forward, the current through the diode is due to
  - 1) majority charge carriers only
  - 2) minority charge carriers only
  - 3) both majority and minority charge carriers
  - 4) none of these
- The depletion region is also called
  - 1) P-region
  - 2) N-region
  - 3) space charge region
  - 4) none
- When avalanche break down occurs, the diode
  - 1) can be used again
  - 2) can not be used again
  - 3) will burn
  - 4) none of these
- When zener break down occurs, the diode
  - 1) can be used again
  - 2) can not be used again
  - 3) will burn
  - 4) none of these
- A diode is a
  - 1) unilateral device
  - 2) bilateral device
  - 3) trilateral device
  - 4) none of these
- Diode is
  - 1) active element
  - 2) passive element
  - 3) linear element
  - 4) none of these



## JLM Special

- Plate current in Triode mainly controlled by
  - 1) grid voltage
  - 2) heater voltage
  - 3) anode to cathode voltage
  - 4) none
- In an electron strikes an atom, and the atom becomes positive ion this process is called
  - 1) ionisation by attraction
  - 2) ionisation by repulsion
  - 3) ionisation by collision
  - 4) none of these
- Zener diode is
  - 1) similar to ordinary diode
  - 2) similar to ordinary diode but it is doped more
  - 3) not similar to ordinary code
  - 4) none of these
- Zener diode is always operated in
  - 1) reverse bias
  - 2) forward bias
  - 3) both reverse and forward bias
  - 4) none of these
- The symbol for zener diode is
  - 1)
  - 2)
  - 3)
  - 4)
- The symbol for SCR is
  - 1)
  - 2)
  - 3)
  - 4)
- When input voltage decrease zener current
  - 1) increase
  - 2) decrease
  - 3) remains same
  - 4) none of these
- Transistor means
  - 1) transfer resistor
  - 2) transmitter
  - 3) transconductor
  - 4) none of these
- Transistor can be used as
  - 1) amplifier
  - 2) switch
  - 3) oscillator
  - 4) all the above
- The number of layers present in transistor?
  - 1) 1
  - 2) 2
  - 3) 3
  - 4) 4
- The doping concentration of collector is
  - 1) less than base
  - 2) more than emitter
  - 3) between base and emitter
  - 4) none of these
- The junction present between emitter and base is called
  - 1) base junction
  - 2) collector junction
  - 3) emitter junction
  - 4) none of these
- The symbol of NPN transistor is
  - 1)
  - 2)
  - 3)
  - 4)

- In a transistor the ratio between collector current and emitter current  $\left(\frac{I_c}{I_e}\right)$  is called
  - 1) transistor  $\alpha$
  - 2) transistor  $\beta$
  - 3) transistor  $\gamma$
  - 4) transistor  $\phi$
- If  $I_e$  is emitter current,  $I_b$  is base current and  $I_c$  is collector current, then
  - 1)  $I_e = I_c - I_b$
  - 2)  $I_e = I_c + I_b$
  - 3)  $I_c = I_e + I_b$
  - 4)  $I_b = I_e + I_c$
- For a transistor ' $\alpha$ ' in terms of ' $\beta$ ' is given by
  - 1)  $\alpha = \frac{1+\beta}{\beta}$
  - 2)  $\alpha = \frac{\beta}{1-\beta}$
  - 3)  $\alpha = \frac{\beta}{1+\beta}$
  - 4)  $\alpha = \frac{1-\beta}{\beta}$
- When the base to emitter is forward biased and collector to base is reverse biased. Transistor is in
  - 1) active region
  - 2) saturation region
  - 3) cut-off
  - 4) none of these
- The semiconductor diode is a
  - 1) unipolar, unilateral device
  - 2) bipolar, unilateral device
  - 3) bipolar, bilateral device
  - 4) unipolar, bilateral device
- Germanium is
  - 1) trivalent
  - 2) tetravalent
  - 3) divalent
  - 4) monovalent
- In a diode tube under the normal working conditions
  - 1) electrons will flow from anode to cathode
  - 2) electrons will flow from cathode to anode
  - 3) electrons will not flow in the tube
  - 4) holes will flow from cathode to anode
- The current gain of CE transistor is
  - 1) zero
  - 2) unity
  - 3) medium
  - 4) high
- For a BJT, under the saturation conditions
  - 1)  $I_c = \beta I_b$
  - 2)  $I_c > \beta I_b$
  - 3)  $I_c < \beta I_b$
  - 4) none of these
- Vacuum triode is
  - 1) current operated device
  - 2) voltage operated device
  - 3) negative resistance device
  - 4) none of these
- BJT is a
  - 1) current operated device
  - 2) voltage operated device
  - 3) negative resistance device
  - 4) none of these
- UJT is
  - 1) current operated device
  - 2) voltage operated device
  - 3) negative resistance device
  - 4) none
- The gate on N-channel JFET is made of
  - 1) P-type semiconductor material
  - 2) N-type semiconductor material
  - 3) silicon dioxide ( $SiO_2$ )
  - 4) deposited metal

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- The gate of N-channel depletion type MOSFET is made of
  - 1) P-type semiconductor material
  - 2) N-type semiconductor material
  - 3) silicon dioxide ( $SiO_2$ )
  - 4) deposited metal
- The gate of N-channel enhancement type MOSFET is made of
  - 1) P-type semiconductor material
  - 2) N-type semiconductor material
  - 3) silicon dioxide ( $SiO_2$ )
  - 4) deposited metal
- Compared to BJT, FET input impedance is
  - 1) low
  - 2) very low
  - 3) high
  - 4) very high
- The two essential conditions to be full filled by the oscillations are
  - 1) generation and amplification
  - 2) amplification and feed back
  - 3) wave shape and feed back
  - 4) generation and feed back
- A crystal oscillator is used when \_\_\_\_\_ frequency is desired
  - 1) low
  - 2) high
  - 3) stable
  - 4) none of these
- SCR is used as switch in
  - 1) low power circuits
  - 2) medium power circuits
  - 3) above two
  - 4) high power circuits
- The switching frequency of transistor is
  - 1) equal to SCR
  - 2) less than SCR
  - 3) more than SCR
  - 4) none of these
- UJT is used to fire
  - 1) SCR
  - 2) FET
  - 3) diode
  - 4) transistor
- Transistor is a \_\_\_\_\_ device
  - 1) Bipolar
  - 2) Unipolar
  - 3) No polar
  - 4) None
- In class A power amplifiers the conducting angle
  - 1)  $360^{\circ}$
  - 2)  $180^{\circ}$
  - 3)  $90^{\circ}$
  - 4)  $270^{\circ}$
- Emitter follower amplifier is generally used as
  - 1) A voltage amplifier
  - 2) A current amplifier
  - 3) A power amplifier
  - 4) A buffer amplifier
- The frequency of oscillation in LC oscillator is
  - 1)  $\frac{1}{2} \pi \sqrt{LC}$
  - 2)  $2\pi \sqrt{2C}$
  - 3)  $2\pi \sqrt{LC}$
  - 4)  $1/2\pi RC$

KEY				
01-1	02-3	03-4	04-3	05-3
06-3	07-1	08-2	09-3	10-3
11-3	12-2	13-1	14-4	15-3
16-1	17-1	18-3	19-2	20-1
21-1	22-1	23-1	24-3	25-2
26-1	27-2	28-4	29-2	30-1
31-4	32-3	33-3	34-3	35-1
36-1	37-2	38-3	39-1	40-2
41-2	42-2	43-3	44-3	45-2
46-1	47-3	48-1	49-3	50-3
51-1	52-2	53-3	54-4	55-2
56-1	57-1	58-1	59-4	60-1

# Aqua regia is a mixture of?

## 17TH GROUP ELEMENTS

Continued from 12<sup>th</sup> November..

15. Halogens have  
1) High ionisation energies  
2) High electronegativities  
3) High electron affinities 4) All the above
16. Which one of the following order is correct for the bond energies?  
1)  $I_2 > Cl_2 > Br_2$  2)  $Br_2 > Cl_2 > I_2$   
3)  $I_2 > Br_2 > Cl_2$  4)  $Cl_2 > Br_2 > I_2$
17. Which halogen requires least energy for excitation  
1) F 2) Cl 3) Br 4) I
18. Boiling point is least for  
1) HF 2) HCl 3) HBr 4) HI
19. Which of the following statements is true  
1)  $Br_2$  is more reactive than  $Cl_2$   
2)  $I_2$  is more reactive than  $Br_2$   
3)  $Cl_2$  is insoluble in water  
4) Iodine is a solid
20. Affinity for hydrogen is maximum for  
1)  $F_2$  2)  $Cl_2$  3)  $Br_2$  4)  $I_2$
21. Which halogen reacts partially with hydrogen even at high temperature and in the presence of a catalyst  
1)  $F_2$  2)  $Cl_2$  3)  $Br_2$  4)  $I_2$
22. Bleaching action of chlorine occurs in the presence of  
1)  $O_2$  2) moisture 3) dry air 4) sunlight
23. The correct order of reactivity is  
1)  $F > Cl > Br > I$  2)  $F < Cl < Br < I$   
3)  $F < Cl > Br > I$  4)  $F > Cl < Br < I$
24. Which of the following reactions is quite violent in nature?  
1) Between hydrogen and fluorine  
2) Between hydrogen and chlorine  
3) Between hydrogen and bromine  
4) Between hydrogen and iodine
25. Perchloric acid is prepared by the action of concentrated sulphuric acid on  
1)  $NaClO_4$  2)  $KClO_3$   
3)  $Ca(ClO_3)_2$  4)  $Ba(ClO_3)_2$
26. The set with correct order of acidity is  
1)  $HClO < HClO_2 < HClO_3 < HClO_4$   
2)  $HClO_4 < HClO_3 < HClO_2 < HClO$   
3)  $HClO < HClO_2 < HClO_3 < HClO_4$   
4)  $HClO_4 < HClO_2 < HClO_3 < HClO$
27. The order of Cl-O bond distance of  $HClO, HClO_2, HClO_3, HClO_4$   
1)  $HClO > HClO_2 > HClO_3 > HClO_4$   
2)  $HClO > HClO_2 > HClO_4 > HClO_3$   
3)  $HClO > HClO_3 > HClO_2 > HClO_4$   
4)  $HClO_4 > HClO_3 > HClO_2 > HClO$
28. Number of sigma and pi bonds in  $ClO_2^-$  ion  
1)  $2\sigma$  and  $2\pi$  2)  $2\sigma$  and  $1\pi$   
3)  $1\sigma$  and  $2\pi$  4)  $3\sigma$  and  $2\pi$
29. Which one of the following oxyacids of chlorine is the least oxidising in nature?  
1) HOCl 2)  $HClO_2$   
3)  $HClO_4$  4)  $HClO_3$
30. Which of the following oxy acids does not contain peroxy bond  
1) Pernitric acid 2) Caro's acid  
3) Perchloric acid 4) Marshall's acid
31. Shape of  $ClO_3^-$  ion is  
1) Pyramidal 2) Tetrahedral  
3) angular 4) linear
32. Number of sigma and pi bonds in  $ClO_3^-$  ion  
1)  $2\sigma$  and  $2\pi$  2)  $3\sigma$  and  $3\pi$   
3)  $3\sigma$  and  $2\pi$  4)  $2\sigma$  and  $3\pi$
33. A powerful oxidant among the following is  
1) Hypochlorite ion 2) Chlorite ion  
3) Chlorate ion 4) Perchlorate ion
34. What is the oxidation state of chlorine in hypochlorous acid?  
1) +1 2) +3 3) +5 4) +7
35. The oxidation number of chlorine in  $KClO_3$  is  
1) +1 2) +3 3) +5 4) +7
36. Fluorine shows anomalous behaviour in VIIA group due to

- 1) its small size 2) high electro negativity  
3) absence of d-orbitals 4) All
37. In the isolation of fluorine a number of difficulties were encountered. Which statement is correct  
1) Electrolysis of aqueous HF gives  $H_2$  and  $O_2$   
2) HF is highly stable and can't be chemically oxidised to  $F_2$   
3) Anhydrous HF is a non conductor of electricity  
4) All
38. Which of the following is the strongest oxidising agent?  
1) F 2) Br 3) Cl 4) I
39. HF present as impurity in gaseous fluorine can be removed by passing over  
1)  $P_2O_5$  2) NaF 3)  $H_2SO_4$  4)  $CaCl_2$
40. Fluorine reacts with water to form  
1) HF and  $O_2$  2) HF and  $O_3$   
3) HF and  $OF_2$  4) HF,  $O_2$  and  $O_3$
41. Fluorine reacts with cold and dil. NaOH and the products are  
1)  $HF + H_2O + NaF$  2)  $NaF + H_2O + OF_2$   
3)  $NaF + H_2O + O_2$  4)  $HF + NaF + O_2$
42. The halogen which can displace three halogens from their compounds is  
1) F 2) Cl 3) Br 4) I
43. Glass reacts with HF forming  
1)  $F_2$  2)  $H_2SiF_6$  3)  $CaF_2$  4) NaF
44. Fluorine does not show variable oxidation states because  
1) it is a halogen 2) of absence of d-orbitals  
3) absence of s-orbital 4) absence of p-orbitals
45. Which one of the following represents the reaction between fluorine and cold dilute NaOH?  
1)  $2F_2 + 4NaOH \rightarrow 4NaF + 2H_2O$   
2)  $3F_2 + NaOH \rightarrow 5NaF + NaFO_3$   
3)  $F_2 + 2NaOH \rightarrow NaF + NaOF + H_2O$   
4)  $2F_2 + 2NaOH \rightarrow 2NaF + OF_2 + H_2O$
46. In the preparation of fluorine, HF is not used because  
1) it is a liquid 2) it is not an electrolyte  
3) it gives different products 4) it is less reactive
47. The compound used in etching of glass is  
1) HF 2)  $HClO_4$  3)  $HNO_3$  4)  $H_2SO_4$
48. Chlorine reacts with caustic soda (hot) giving NaCl and  
1) NaClO 2)  $NaClO_2$  3)  $NaClO_3$  4)  $NaClO_4$
49. In Nelson's cell, Cathode is  
1. Graphite 2) Copper tube  
3. Porus steel vessel 4) Iron tank
50. What are the products formed when ammonia reacts with excess chlorine?  
1)  $N_2$  and  $NCl_3$  2)  $NCl_3$  and HCl  
3)  $N_2$  and  $NH_4Cl$  4)  $N_2$  and HCl
51. DDT is  
1) Di phenyl trichloro ethene  
2) Di chloro di phenyl triethane  
3) Di chloro difluoro tri phenyl ethene  
4) Di chloro diphenyl tri chloro ethane
52. The reaction of chlorine with CO in the presence of sun light gives  
1)  $COCl_2$  2)  $CO_2Cl_2$  3) HOCl 4)  $H_2Cl_2O_2$
53. The chemical formula of tear gas is  
1)  $COCl_2$  2)  $CO_2$  3)  $Cl_2$  4)  $CCl_3NO_2$

### LEVEL-I B KEY

- 15) 4 16) 4 17) 4 18) 2 19) 4 20) 1 21) 4  
22) 2 23) 1 24) 1 25) 1 26) 1 27) 1 28) 2  
29) 3 30) 3 31) 1 32) 3 33) 1 34) 1 35) 3  
36) 4 37) 4 38) 1 39) 2 40) 4 41) 2 42) 1  
43) 2 44) 2 45) 4 46) 2 47) 1 48) 3 49) 3  
50) 2 51) 4 52) 1 53) 4

### LEVEL II A

#### GENERAL CHARACTERISTICS

1. What is the formal charge on Cl in the following lewis structure
- $$\begin{array}{c} \text{:}\ddot{\text{O}}\text{:} \\ | \\ \text{:}\ddot{\text{O}}\text{---Cl---}\ddot{\text{O}}\text{---H} \\ | \\ \text{:}\ddot{\text{O}}\text{:} \end{array}$$
- 1) 0 2) +3 3) +5 4) +7
2. All though electron gain enthalpy of fluorine is less negative as compared to chlorine, Fluorine is stronger oxidizing agent than Chlorine because of  
i) Low Enthalpy of dissociation of F-F bond

- ii) High Hydration Enthalpy of  $F^-$   
iii) High Electron affinity of fluorine than chlorine

- 1) i,ii correct 2) i,iii correct  
3) i,ii,iii correct 4) ii,iii correct
3. Among C-X bond (where X =Cl, Br, I) the correct decreasing bond energy order is  
1)  $C-Cl > C-Br > C-I$  2)  $C-I > C-Br > C-Cl$   
3)  $C-Br > C-Cl > C-I$  4)  $C-I < C-Br < C-Cl$
4. The maximum negative electron gain enthalpy of halogens in their respective periods is because of  
1) Small size in their respective period  
2) High effective nuclear charge  
3) More electro negativity in their respective periods  
4) All the above
5.  $IBr_3$  cannot exist but  $IF_3$  exist. This fact can be explained on the basis of  
1) Electro -vities 2) Electron affinities  
3) Ratio of radii of atoms 4) Reducing abilities

#### OXYACIDS & HALOGEN PREPARATION

6. The increasing order of the acidity of the following acids : I= $H_2SO_4$ , II= $H_3PO_3$  and III= $HClO_3$   
1) I>II>III 2) II<I<III 3) III>I>II 4) II>III>I
7. Among the oxyacids of chlorine, the strongest oxidising agent  
1)  $HClO_4$  2)  $HClO_3$  3)  $HClO_2$  4)  $HClO$

### JEE MAIN Special

8. Find the correct statements.  
a) Electron affinity of F is less than that of Cl  
b) Number of lone pairs at central chlorine atom of  $ClF_3$  is 2  
c) Iodine absorbs radiation of violet colour and appear in yellow colour.  
d)  $F_2$  oxidizes all other ionic halides to halogens  
Find the correct answer.  
1) a, c, d 2) a, b, d 3) only c 4) all are correct
9. In Moissan's method the electrodes were  
1) Graphite 2) Pt  
3) Ni 4) Pt-Ir alloy
10. In Moissan's method Fluorine was prepared by the electrolysis of  
1) Fused HF 2) Molten KF  
3) HF dissolved in Molten KF 4)  $OF_2$
11. Iodine can be obtained from NaI solution by the action of  
1) Chlorine & bromine 2) Fluorine  
3) Soluble chloride 4) Soluble bromine
12. Aqua regia is a mixture of?  
1) 3 : 1 ratio of conc HCl and Conc.  $HNO_3$   
2) 3 : 1 ratio of conc  $HNO_3$  and Conc. HCl  
3) 1 : 3 ratio of conc HCl and Conc.  $HNO_3$   
4) 1 : 3 ratio of conc HCl and Conc.  $H_2SO_4$
13. Following is incorrect statement  
1) Bond dissociation energy of Bromine is more than that of chlorine.  
2) The standard electrode potential of chlorine is more than that of fluorine  
3) In bleaching powder, the theoretical percentage of chlorine is 35-38%.  
4) All

#### PREVIOUS EAMCET QUESTIONS

14. Which of the following reactions does not occur? (2008 E)  
1)  $F_2 + 2Cl^- \rightarrow 2F^- + Cl_2$   
2)  $Cl_2 + 2F^- \rightarrow 2Cl^- + F_2$   
3)  $Br_2 + 2I^- \rightarrow 2Br^- + I_2$   
4)  $Cl_2 + 2Br^- \rightarrow 2Cl^- + Br_2$
15. Identify B in the above reaction.  
 $Ca(OH)_2 \xrightarrow{Cl_2, -H_2O} A$   
 $\xrightarrow{\text{Auto Oxidation}} CaCl_2 + B$   
1)  $CaOCl_2$  2)  $Ca(ClO_3)_2$   
3)  $Ca(OH)_2$  4)  $Ca(ClO_2)_2$

# విజేత

For Feedback...

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16. Fluorine reacts with dilute NaOH and forms a gaseous product A. The bond angle in the molecule of A is  
1)  $104^\circ 40'$  2)  $103^\circ$  3)  $107^\circ$  4)  $109^\circ 28'$
17. Assertion(A) : The bond dissociation energy of fluorine is less than bromine.  
Reason (R) : In fluorine molecule, large lone pair electronic repulsions and appreciable inter nuclear repulsions are present. (2009 M)  
1) Both A and B are true and R is the correct explanation of A.  
2) Both A and R are true but R is not the correct explanation of A.  
3) A is true but R is not true  
4) A is not true but R is true
18. Which one of the following is formed apart from sodium chloride when chlorine reacts with hot concentrated sodium hydroxide?  
1)  $NaOCl$  2)  $NaClO_3$  3)  $NaClO_2$  4)  $NaClO_4$
19. The order of bond energies in halogen molecules is  
1)  $F_2 < Cl_2 < Br_2 < I_2$  2)  $F_2 > Cl_2 > Br_2 > I_2$   
3)  $Cl_2 > Br_2 > F_2 > I_2$  4)  $Cl_2 > F_2 > Br_2 > I_2$
20. Bond energy of  $Cl_2, Br_2$  and  $I_2$  follow the order  
1)  $Cl_2 > Br_2 > I_2$  2)  $Br_2 > Cl_2 > I_2$   
3)  $I_2 > Br_2 > Cl_2$  4)  $I_2 > Cl_2 > Br_2$

#### LEVEL-II A - KEY

- 1) 2 2) 1 3) 1 4) 4 5) 3 6) 2 7) 4  
8) 2 9) 4 10) 3 11) 1 12) 113) 4 14) 2  
15) 2 16) 2 17) 1 18) 2 19) 3 20) 1

#### LEVEL II B

#### GENERAL CHARACTERISTICS

1. Which of the following is not true of the halogens  
1) All of them have seven electrons in their outer shell  
2) They either gain an electron by forming an ionic bond or form a covalent compound by electron sharing  
3) All exhibit variable valency  
4) All form diatomic molecules

#### OXIDATION STATES

2. Which requires highest energy for excitation  
1) F 2) Cl 3) Br 4) I

#### PHYSICAL STATE

3. Halogen elements are those which are  
1) diatomic and form  $X^-$  ions  
2) monoatomic and form  $X^-$  ions  
3) diatomic and form  $X^-$  ions  
4) monoatomic and form  $X^-$  ions

#### BOND DISSOCIATION ENERGY

4. The decreasing order of bond dissociation energy among halogens is  
1)  $F > Cl > Br > I$   
2)  $Cl > Br > F > I$   
3)  $Cl > Br > I > F$   
4)  $I > F > Cl > Br$

#### LEVEL-II B - KEY

- 1) 3 2) 1 3) 3 4) 2