



1 1.008 H hydrogen																	2 4.003 He helium															
3 6.94 Li lithium	4 9.012 Be beryllium											5 10.81 B boron	6 12.01 C carbon	7 14.01 N nitrogen	8 16.00 O oxygen	9 19.00 F fluorine	10 20.18 Ne neon															
11 22.99 Na sodium	12 24.31 Mg magnesium											13 26.98 Al aluminium	14 28.09 Si silicon	15 30.97 P phosphorus	16 32.06 S sulfur	17 35.45 Cl chlorine	18 39.95 Ar argon															
19 39.10 K potassium	20 40.08 Ca calcium	21 44.96 Sc scandium	22 47.87 Ti titanium	23 50.94 V vanadium	24 52.00 Cr chromium	25 54.94 Mn manganese	26 55.85 Fe iron	27 58.93 Co cobalt	28 58.69 Ni nickel	29 63.55 Cu copper	30 65.38 Zn zinc	31 69.72 Ga gallium	32 72.63 Ge germanium	33 74.92 As arsenic	34 78.96 Se selenium	35 79.90 Br bromine	36 83.80 Kr krypton															
37 85.47 Rb rubidium	38 87.62 Sr strontium	39 88.91 Y yttrium	40 91.22 Zr zirconium	41 92.91 Nb niobium	42 95.96 Mo molybdenum	43 [98] Tc technetium	44 101.1 Ru ruthenium	45 102.9 Rh rhodium	46 106.4 Pd palladium	47 107.9 Ag silver	48 112.4 Cd cadmium	49 114.8 In indium	50 118.7 Sn tin	51 121.8 Sb antimony	52 127.6 Te tellurium	53 126.9 I iodine	54 131.3 Xe xenon															
55 132.9 Cs caesium	56 137.3 Ba barium											72 178.5 Hf hafnium	73 180.9 Ta tantalum	74 183.8 W tungsten	75 186.2 Re rhenium	76 190.2 Os osmium	77 192.2 Ir iridium	78 195.1 Pt platinum	79 197.0 Au gold	80 200.6 Hg mercury	81 204.4 Tl thallium	82 207.2 Pb lead	83 209.0 Bi bismuth	84 [209] Po polonium	85 [210] At astatine	86 [222] Rn radon						
87 [223] Fr francium	88 [226] Ra radium											104 [261] Rf rutherfordium	105 [268] Db dubnium	106 [269] Sg seaborgium	107 [270] Bh bohrium	108 [269] Hs hassium	109 [278] Mt meitnerium	110 [281] Ds darmstadtium	111 [281] Rg roentgenium	112 [285] Cn copernicium	113 [286] Uut ununtrium	114 [289] Ff flerovium	115 [288] Uup ununpentium	116 [293] Lv livermorium	117 [294] Uus ununseptium	118 [294] Uuo ununoctium						
																		57 138.9 La lanthanum	58 140.1 Ce cerium	59 140.9 Pr praseodymium	60 144.2 Nd neodymium	61 [145] Pm promethium	62 150.4 Sm samarium	63 152.0 Eu europium	64 157.3 Gd gadolinium	65 158.9 Tb terbium	66 162.5 Dy dysprosium	67 164.9 Ho holmium	68 167.3 Er erbium	69 168.9 Tm thulium	70 173.1 Yb ytterbium	71 175.0 Lu lutetium
																		89 [227] Ac actinium	90 232.0 Th thorium	91 231.0 Pa protactinium	92 238.0 U uranium	93 [237] Np neptunium	94 [244] Pu plutonium	95 [243] Am americium	96 [247] Cm curium	97 [247] Bk berkelium	98 [251] Cf californium	99 [252] Es einsteinium	100 [257] Fm fermium	101 [258] Md mendelevium	102 [259] No nobelium	103 [262] Lr lawrencium

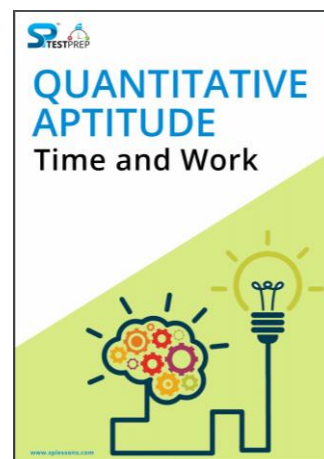
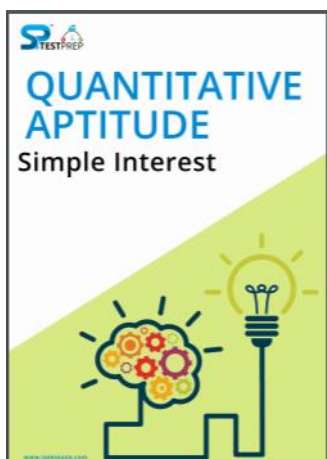
PERIODIC TABLE MOST ASKED QUESTION IN RRB EXAMS

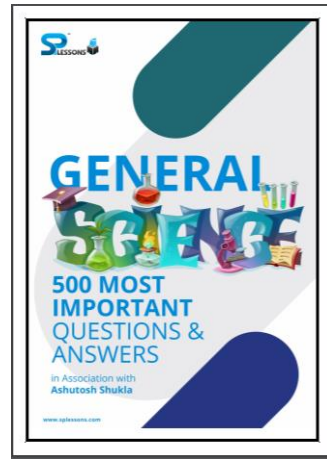
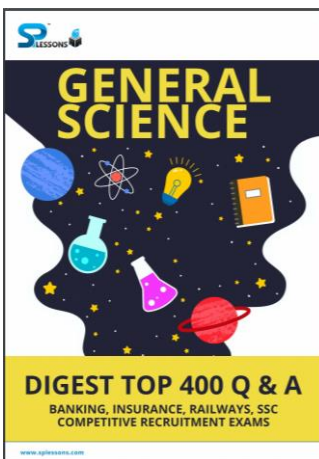
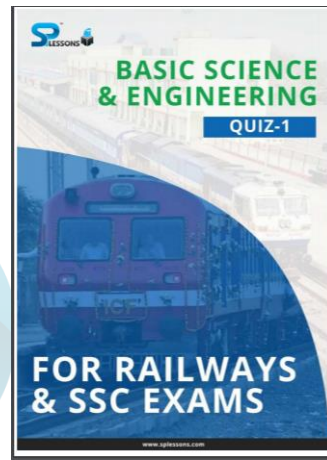
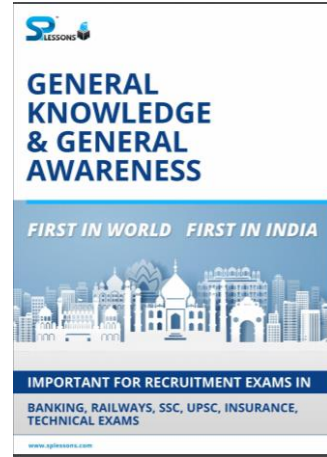


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PERIODIC TABLE MOST ASKED QUESTION IN RRB EXAMS

<u>S.No</u>	<u>Questions</u>	<u>Answer</u>
1.	Most abundant element on earth's crust	Oxygen
2.	Most abundant metal in earth's crust	Aluminum
3.	Most abundant metalloid in earth's crust	Silicon
4.	Most abundant element in atmosphere	Nitrogen
5.	Most abundant element in the universe	Hydrogen
6.	Most abundant element in human body	Oxygen
7.	Most abundant gas in atmosphere	Nitrogen
8.	Most abundant element in sea water	TM Chlorine
9.	Most abundant element in moon's surface	Titanium
10.	Most abundant metal present in human body and bones	Calcium
11.	Most abundant metal compound in bones	Calcium phosphate
12.	Most abundant compound on earth's surface	Water (H ₂ O)
13.	Most abundant compound in sea water	Sodium Chloride
14.	Second most abundant compound in sea water	Magnesium Chloride
15.	Most chemically reactive element	Fluorine



16.	Second most chemically reactive element	Chlorine
17.	The lightest and simplest element	Hydrogen
18.	The lightest metal	Lithium
19.	The rarest element in the earth	Astatine
20.	The heaviest element	Osmium
21.	The heaviest gaseous element	Radon
22.	First man made element	Technetium
23.	Most stable element	Lead
24.	The periodic table is divided into 4 main blocks	*S block - elements of group 1 and 2 *P block - elements of group 13 to 18 *D block - elements of group 3 to 12 *F block - Lanthanides and Actinides
25.	The shortest period of the periodic table	First period
26.	The longest period of the periodic table	Sixth period
27.	Elements with atomic number 57-71 are known as	Lanthanides (rare earths)
28.	Elements with atomic number 89-103 are known as	Actinides (radioactive rare earths)
29.	Total number of elements in periodic table	118
30.	The number of naturally occurring elements	92



31.	Which year is declared as International Year of Periodic Table of Chemical Elements to commemorate the 150th birth anniversary of periodic table of chemical elements?	2019
32.	How many rare earth elements are present in the periodic table?	17
33.	What was the position of noble gases in the Mendeleev's periodic table?	Noble gases were not discovered by that time
34.	Electronegativity _____ from left to right in a row in the periodic table and electro positivity _____ from left to right in a row in the periodic table	increases, decreases
35.	Elements in the Modern Periodic Table are arranged in _____ periods.	7
36.	In the modern periodic table. Periods _____ and _____ contain Lanthanoides and Actinoides.	6 and 7
37.	An element of atomic number 16 will belong to the period of the periodic table.	3rd
38.	How many groups and periods are present in the Modern Periodic Table?	18 groups and 7 periods
39.	Which of the following inert gases is placed in period 4 of the periodic table?	Kr
40.	The only block which consists of metals, non-metals and metalloids in a periodic table is?	P-Block
41.	How many periods are there in a modern periodic table?	7
42.	Who predicted the existence of some yet to be discovered elements on the basis of gaps in his periodic table?	Mendeleev
43.	Who has developed the Modern Periodic table?	Henry Moseley
44.	Picogens are arranged in which block of the periodic table?	p-block
45.	_____ is the only non-metal present in Group 1 of the Modern Periodic Table.	Hydrogen



46.	If the elements are arranged in order of increasing atomic masses, every eight elements has properties similar to the first is known as _____	Newland's law of octaves
47.	Lothar Mayer plotted a graph between atomic volume of the elements and their _____	Atomic mass
48.	In Mendeleev's periodic table he arranged the elements in the _____ atomic weights.	Increasing
49.	Periodic table consists of seven horizontal rows called _____	Periods
50.	Periodic table consists of nine vertical columns called as _____	Groups
51.	The group number of an element represents its _____	Valency
52.	The zero groups contain _____ gases.	Noble
53.	Mendeleev's periodic table Defects?	Systematic study of the elements
54.	The properties of elements are periodic functions of their atomic numbers may be stated as _____	The modern periodic law
55.	How many groups are there in the periodic table?	18
56.	Both of elements of 1st period contains valence electrons in _____	K shell
57.	In periodic table, helium is placed at _____	Top right corner
58.	Across period atomic size decreases due to _____	Increase in nuclear force of attraction
59.	First three periods are _____	Short periods
60.	Chemical properties depends upon _____	Valence shell electronic configuration



61.	Physical properties depends on the	Size of atom
62.	Elements that lie in same column have	Similar properties
63.	As we go from left to right across period, electron affinity	Increases
64.	Elements are arranged in order of	Increasing atomic number
65.	Decrease in force of attraction between valence electrons and nucleus by inner electrons is called	Shielding effect
66.	(Metalloids which resemble metals as well as non-metals are on the border line of metals and non-metals arranged in a _____manner.	zig-zag
67.	There are 118 elements discovered so far, which are classified into	18 groups and 7 periods.
68.	Elements with atomic numbers 58 to 71 are called	lanthanoids
69.	Elements with atomic numbers 90 to 103 are called	actinoids
70.	Properties of elements depend upon the number of _____	valence electrons
71.	Over 80% of the elements are metals are remaining are	non-metals
72.	Metals and non-metals differ in their _____	properties
73.	Non-metals are good oxidising agents, because they can gain_____	electrons easily
74.	How are the elements arranged in the Periodic Table ?	In order of increasing atomic number.
75.	An element belongs to which group in the periodic table if it has an electronic configuration of 2, 8, 7?	Group 17



76.	An element with electronic configuration 2, 8, 8 belongs to which group?	Group 18 (0)
77.	Heaviest element in the periodic table?	Uranium
78.	How many elements are there in the 4th group of the periodic table?	4
79.	How many metalloids are there in the Periodic table?	6
80.	How many periods and groups are there in the Modern Periodic Table?	7 periods and 18 groups
81.	How many Radioactive elements are there in the Periodic Table?	38
82.	If we go from left to right in the periodic table, the capacity to lose electrons will__	Decrease
83.	In the periodic table, to which block does the elements from atomic number 57-72 belongs?	d-block
84.	In which group do Halogens belong in the periodic table?	Group 17
85.	In which of the Group 18 elements, 8 electrons are not present in the outermost shell?	Helium
86.	No. of elements in the 18th group of the Periodic table?	6
87.	The elements 19 to 36 come under which period of the periodic table?	4th period
88.	To which group do elements having electronic configuration 2, 8, 3 belong?	Group 13
89.	What are the elements in the last group of periodic table called?	Noble gases
90.	What does the zig-zag line in Periodic Table imply?	The line which separates the Metallic and Non-metallic elements



91.	What is the Atomic number of Cobalt?	27
92.	What is the Atomic Number of Potassium in the Modern Periodic Table?	19
93.	What is the Valancy of elements in the 16th Group of the periodic table	2
94.	Which element has least atomic mass?	Hydrogen
95.	Which is the last element in the Newland Octave periodic Table	Thorium
96.	Which is the longest period in the periodic table?	6th
97.	Which period has 8 elements in the periodic table?	2nd and 3rd Period
98.	Who invented Periodic Table	Dmitri Mendeleev
99.	Which two elements don't fit according to the Mendeleev as well as Newland table?	Nickle and cobalt
100.	Which is the smallest periodic row	1st



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